

**B.Sc. Semester-VI
Group-A / DSE-4
Organic Synthesis**

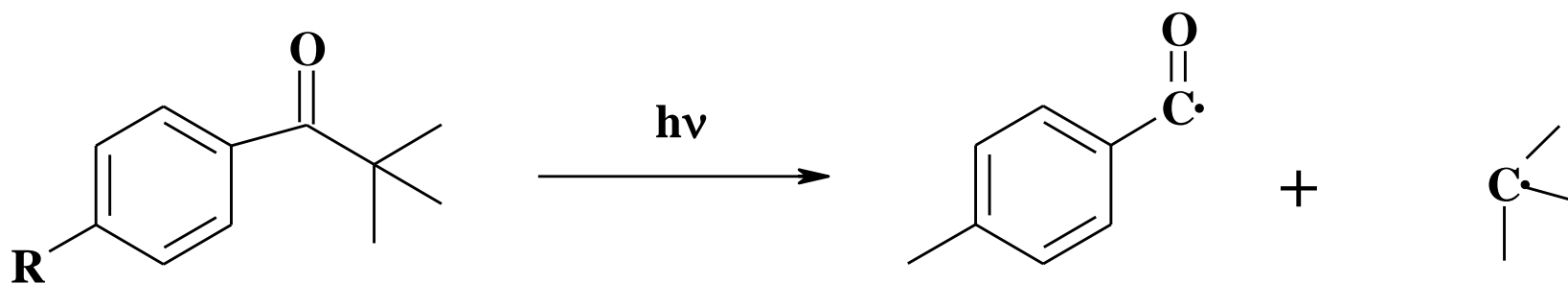


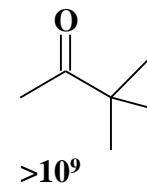
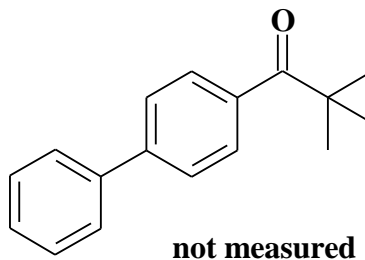
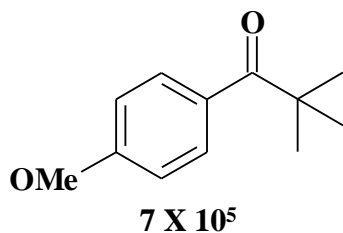
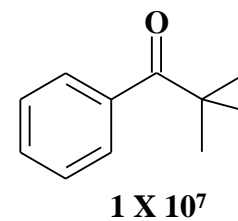
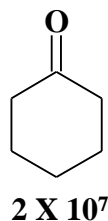
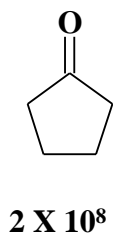
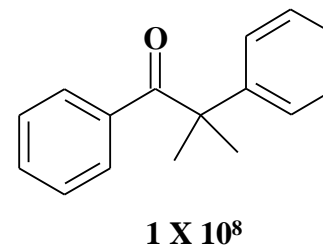
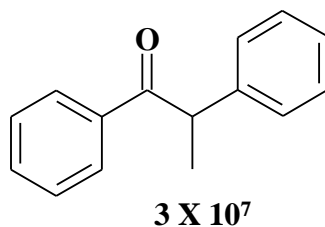
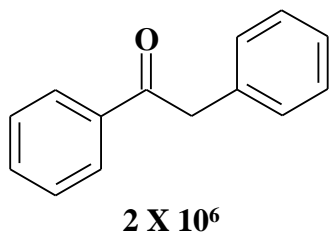
**III. Photochemistry
6. Norrish Type I Reaction**



**Dr. Rajeev Ranjan
University Department of Chemistry
Dr. Shyama Prasad Mukherjee University, Ranchi**

Norrish Type I Reaction



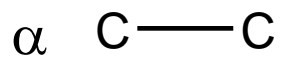
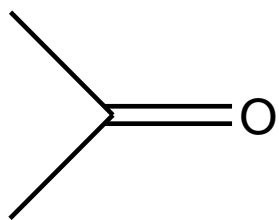
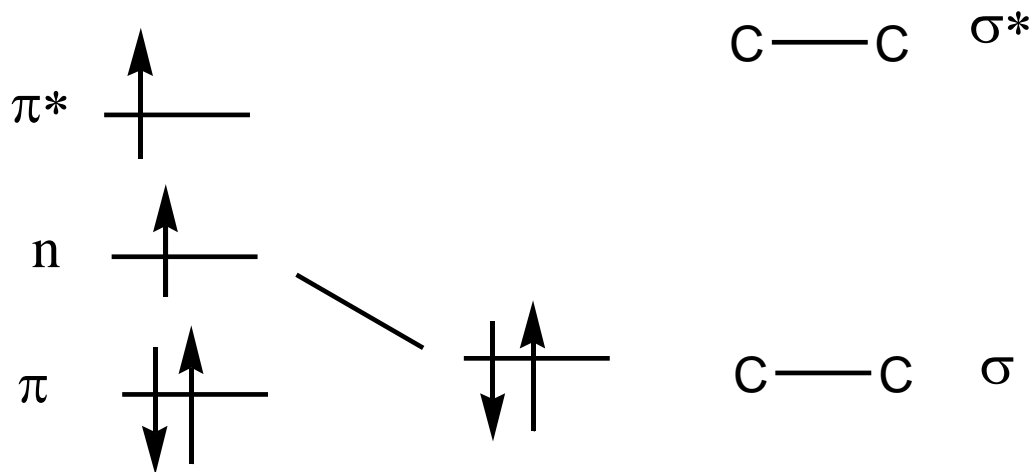
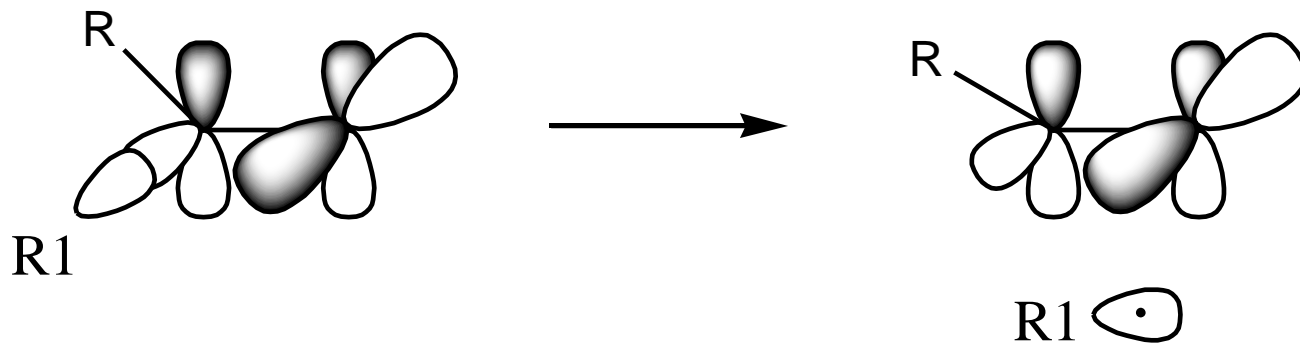


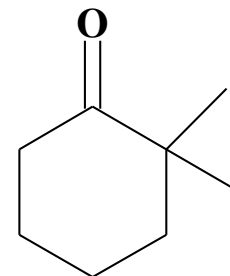
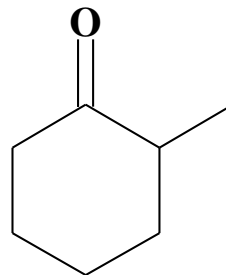
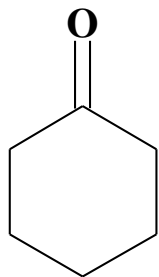
Norrish type I reaction is much faster for $n-\pi^*$ compared to $\pi-\pi^*$ excited states

$n-\pi^*$ reactivity is due to the weakening of the α -bond by overlap of this bond with the half vacant n-orbital of oxygen.

This overlap is not possible for $\pi-\pi^*$ excited states

Electron releasing group at para position lead to stabilization of $\pi-\pi^*$ excited states hence decrease in reactivity



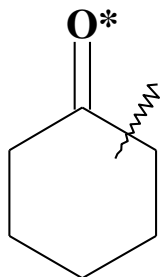


K_{α}

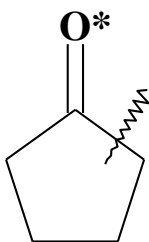
$3.3 \times 10^7/s$

$4.7 \times 10^8/s$

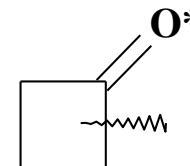
$1.8 \times 10^9/S$



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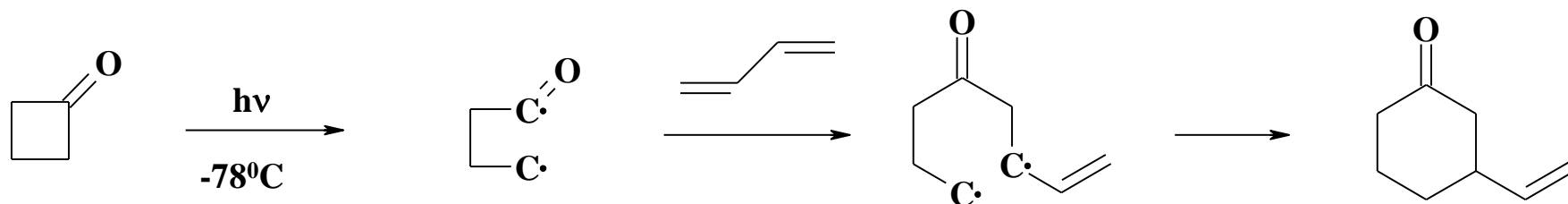
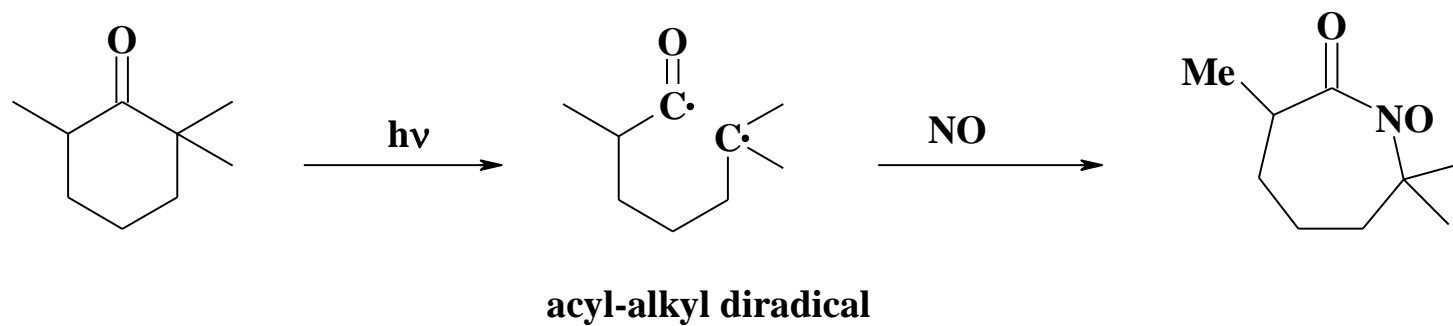


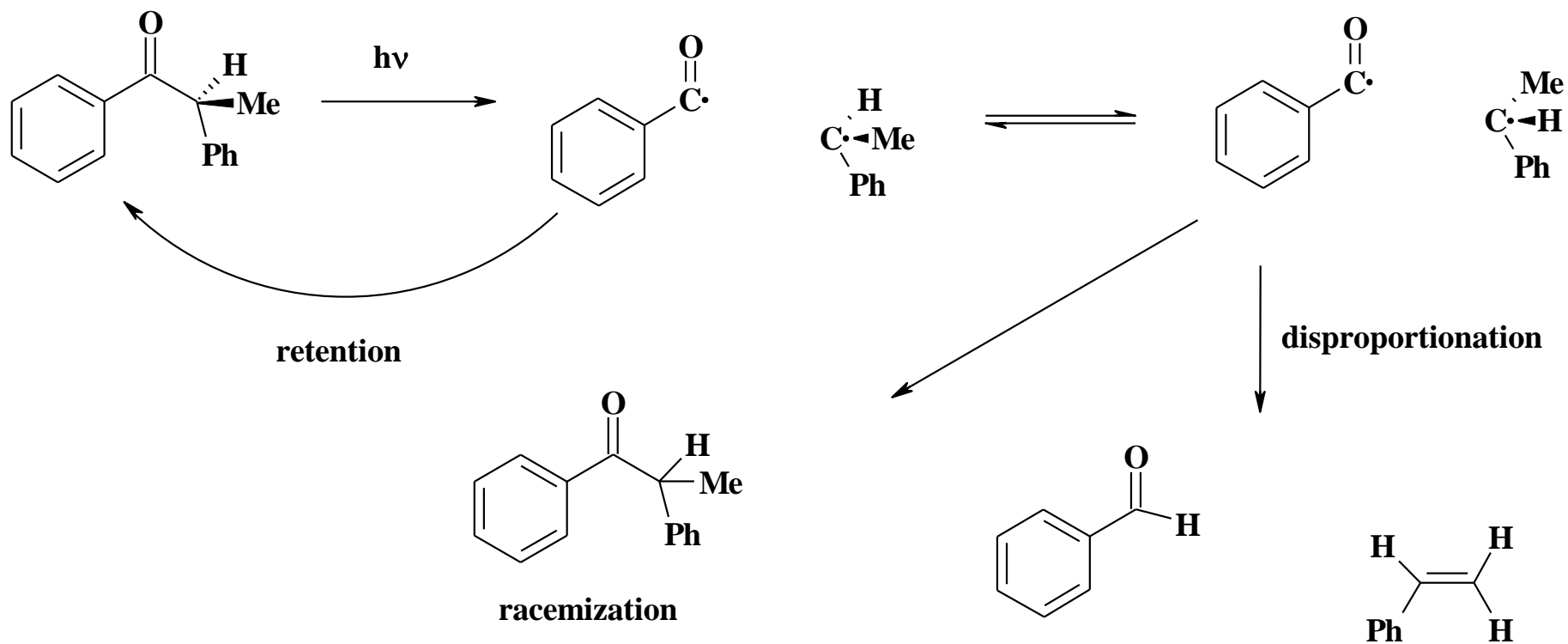
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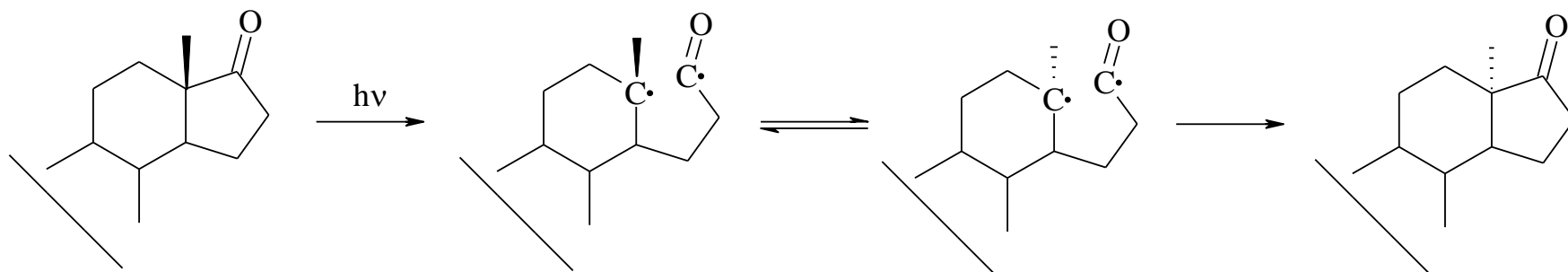
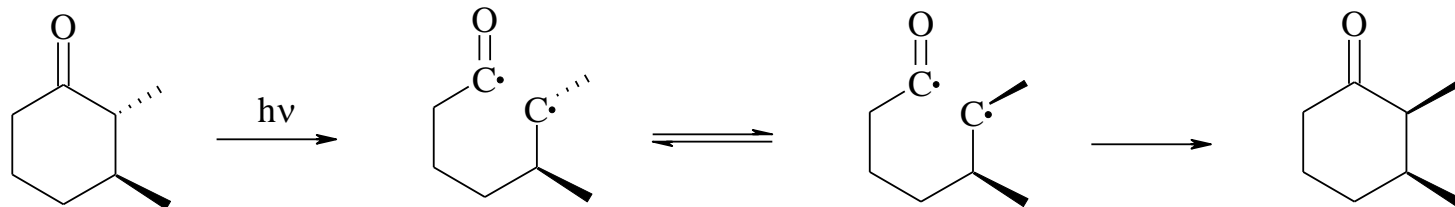


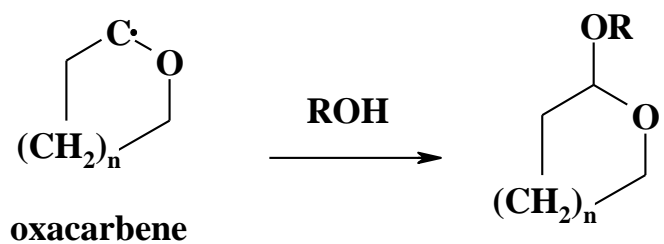
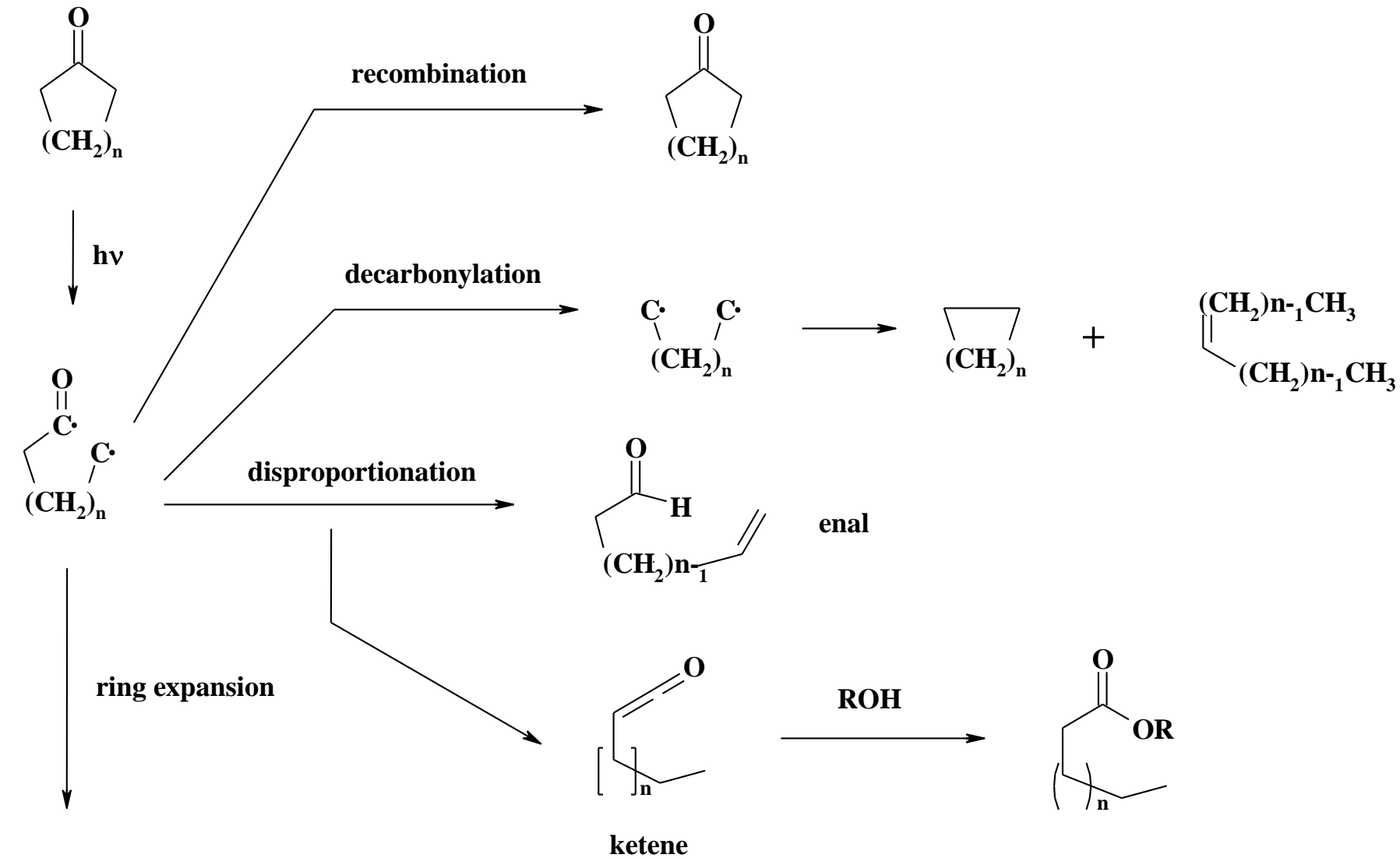
Rate of α cleavage increasing ring strain

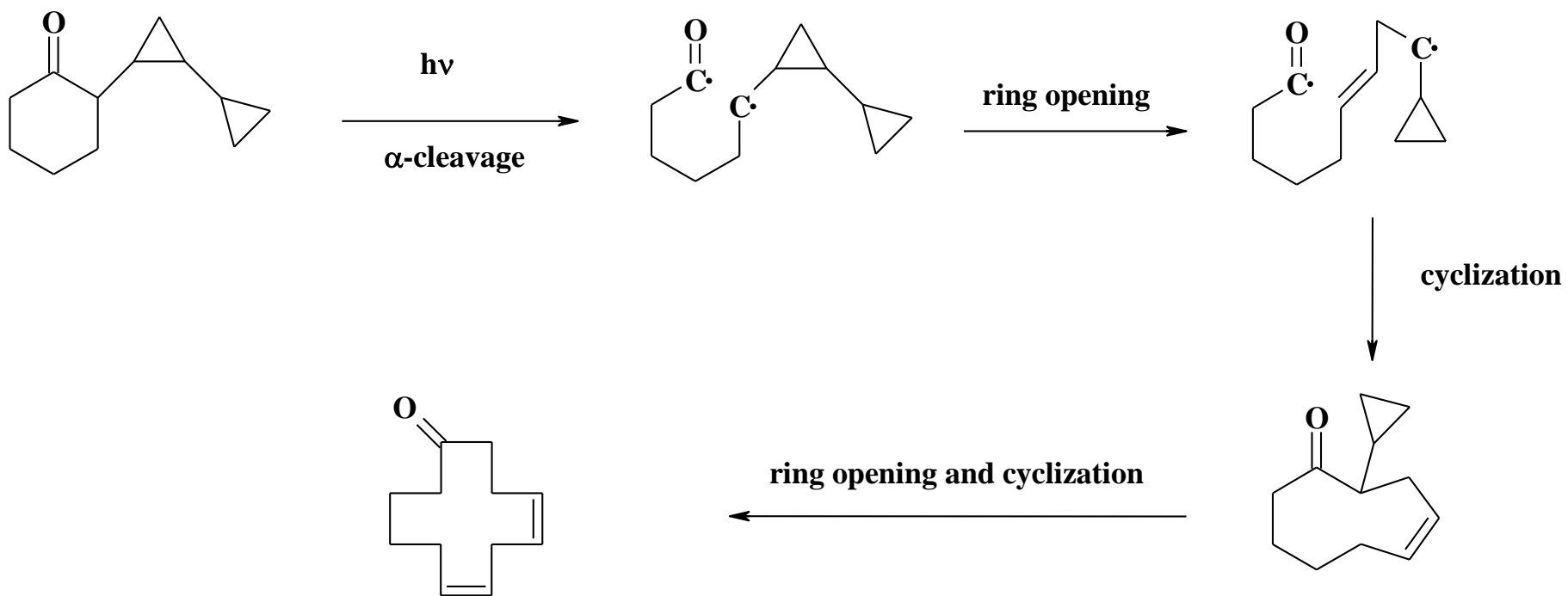
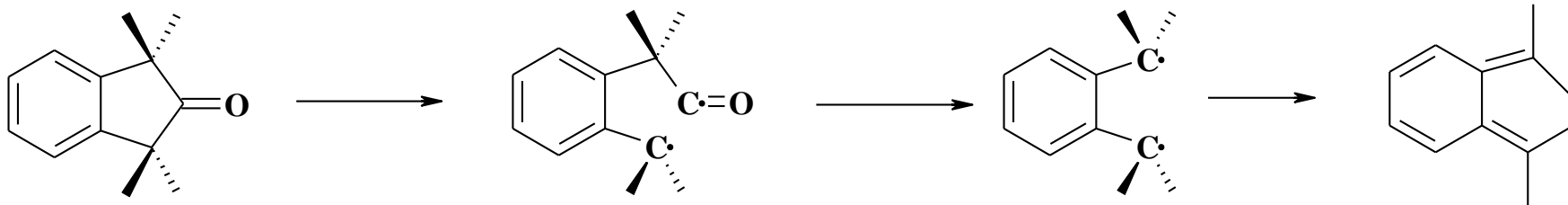
Intermediate trapping experiment

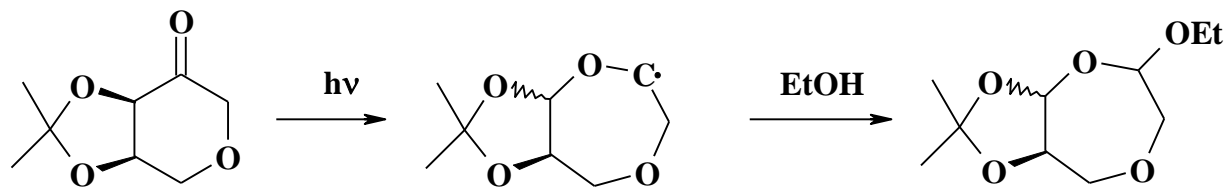
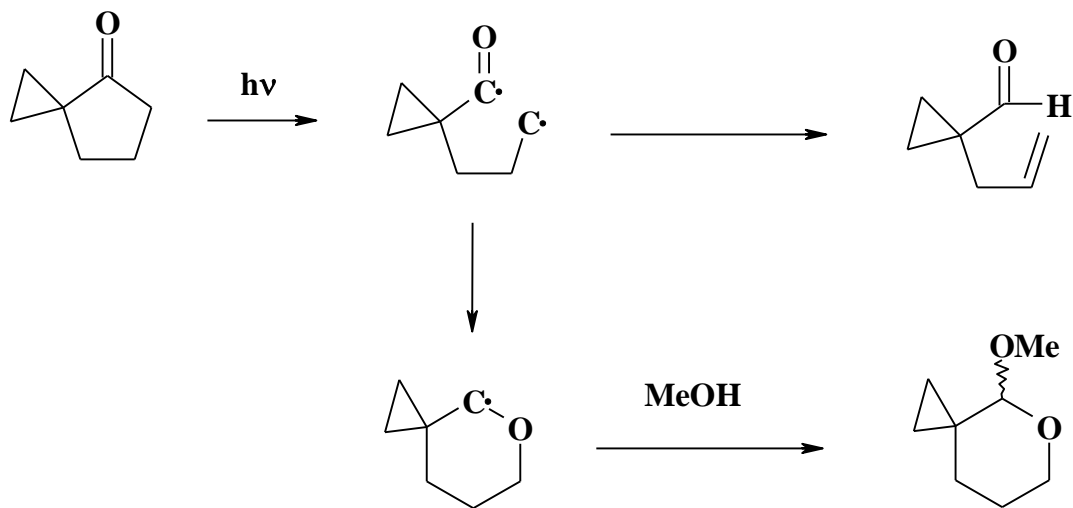
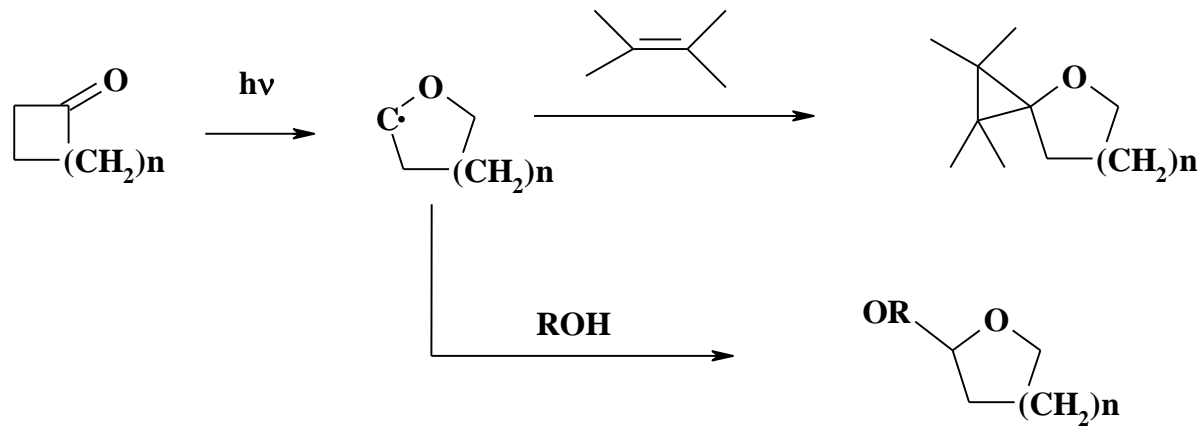


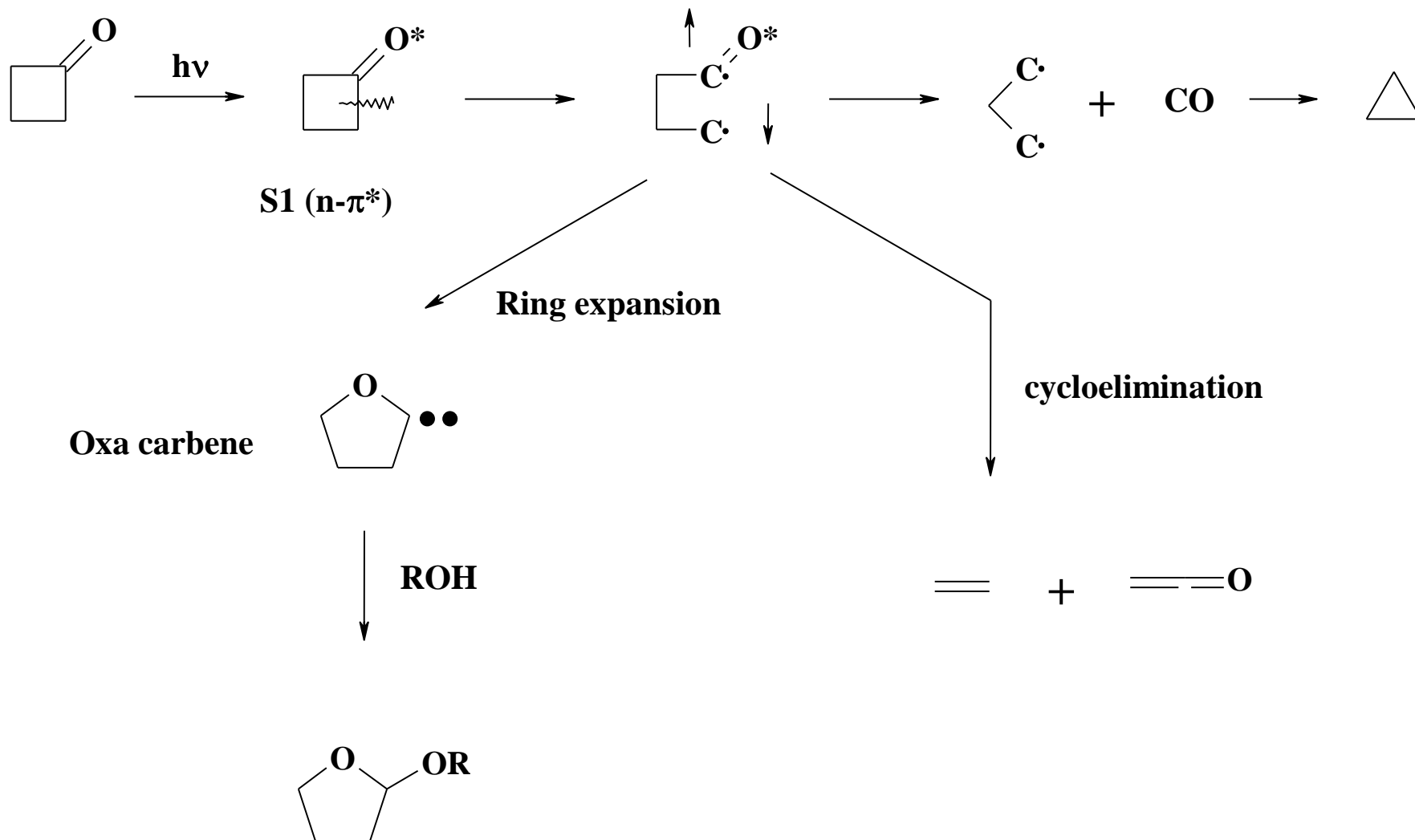


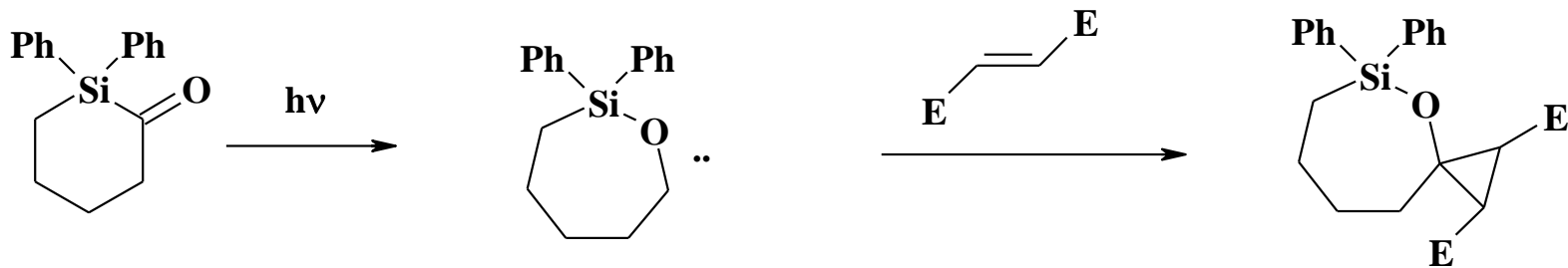
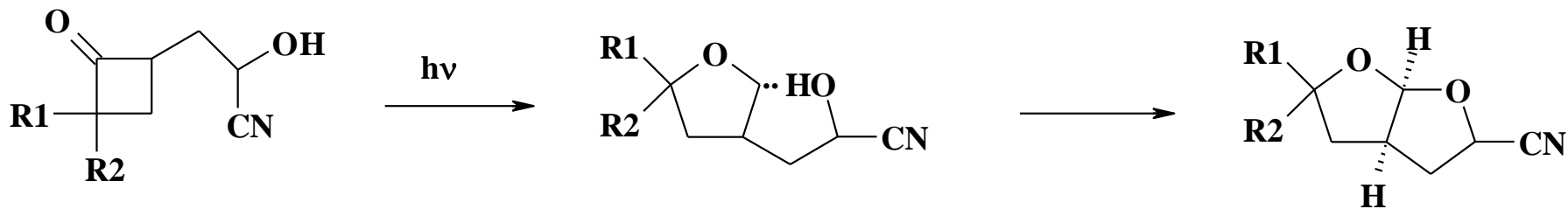
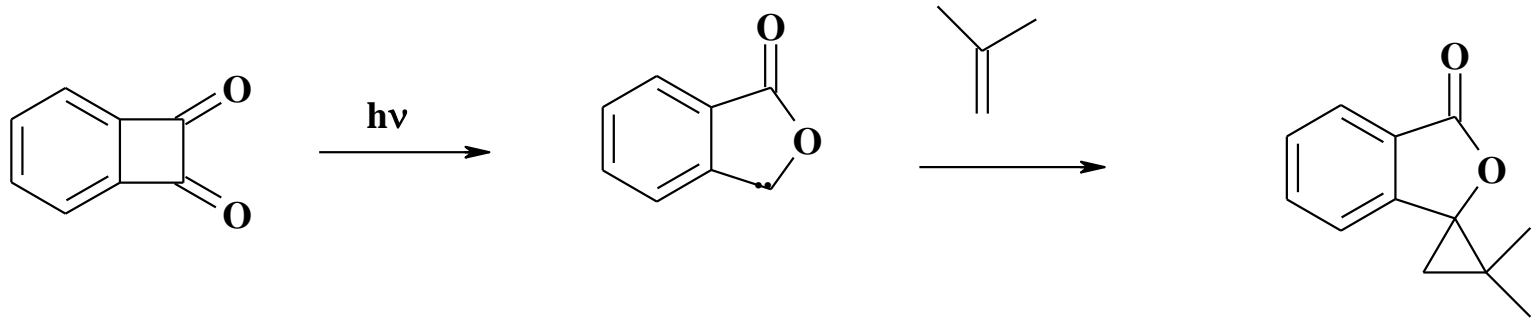


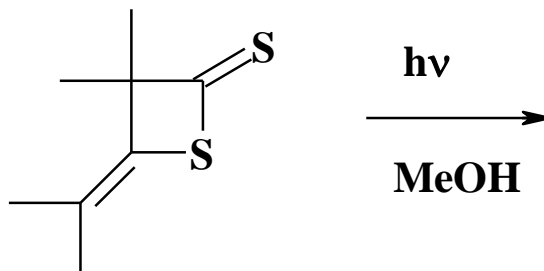
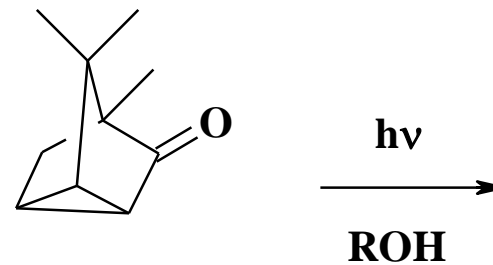
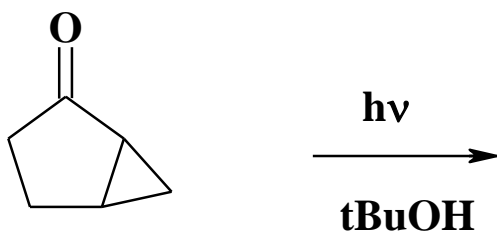
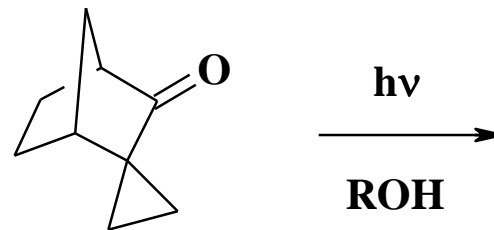
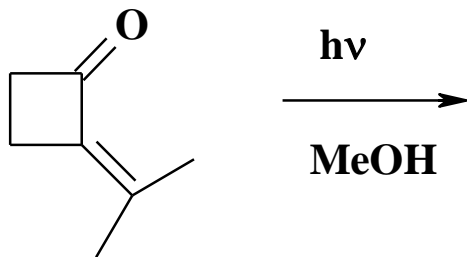


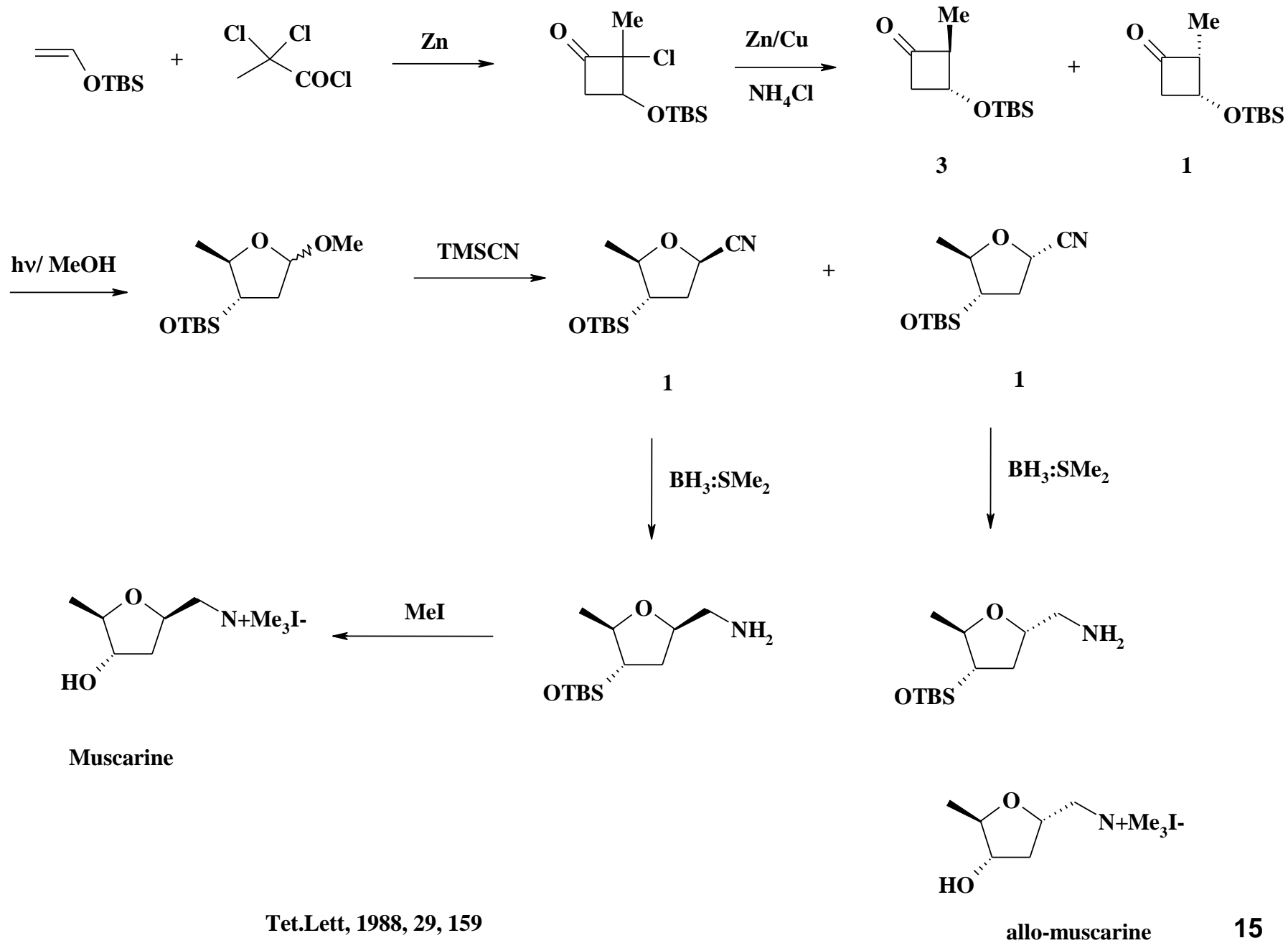


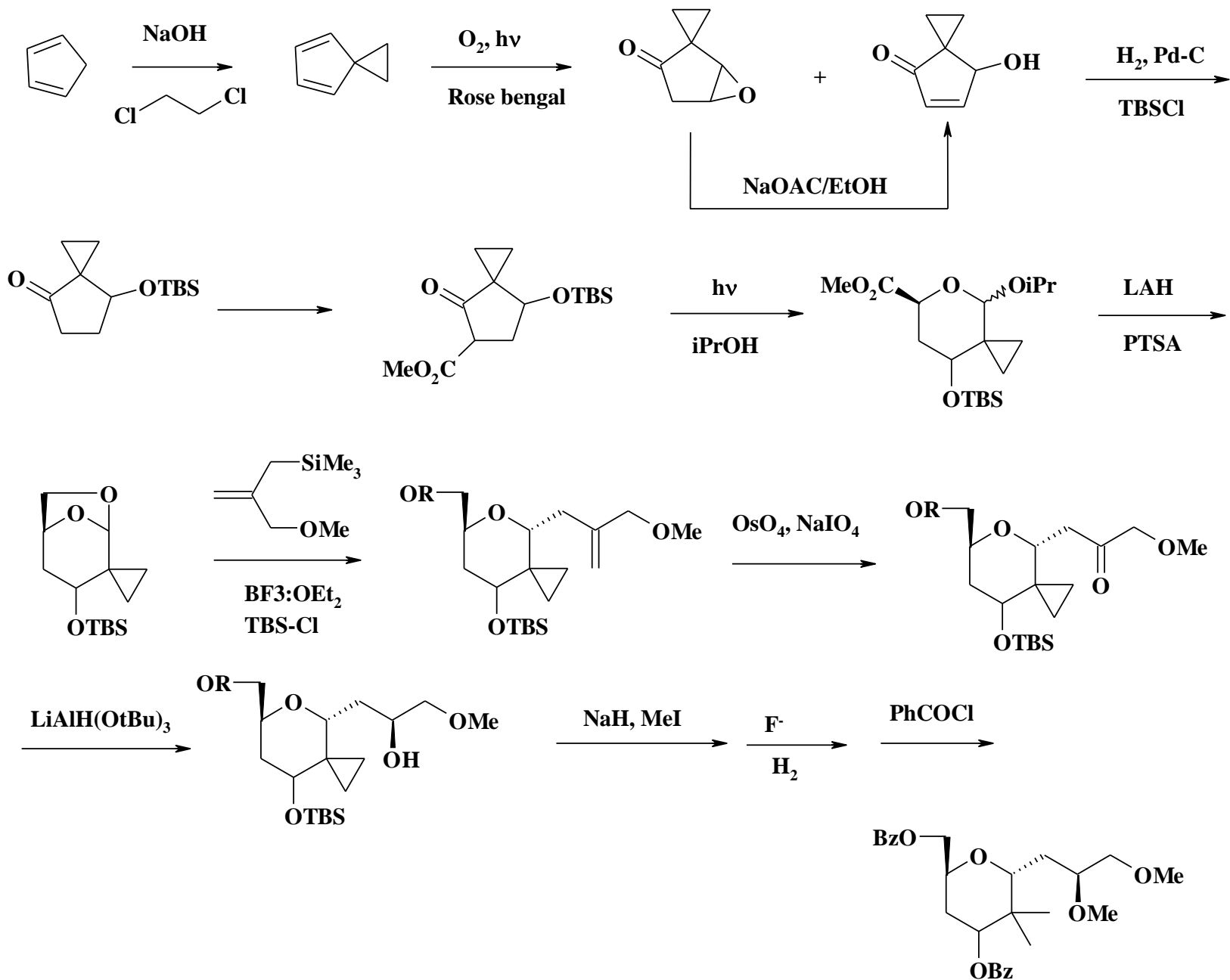












J. Org. Chem. 1987, 52, 2335

Pederol dibenzoate

Thank You



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